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will precipitate  
when 150. mL of  
0.500 M silver  
nitrate are

added to 100. mL  
of 0.400 M

potassium

chromate?  $2 \text{ AgNO}_3$

$3(\text{aq}) + \text{K}_2\text{CrO}_4$

$4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4$

$4(\text{s}) + 2 \text{ KNO}_3$

$3(\text{aq}) \quad 0.150 \text{ L}$

$\text{AgNO}_3 \quad 0.500$

moles  $\text{AgNO}_3 \quad 1$



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moles  $\text{Ag}_2\text{CrO}_4$  4

331.74 g  $\text{Ag}_2$

$\text{CrO}_4 = 12.4$  g

$\text{Ag}_2\text{CrO}_4$  1 L 2

moles ...

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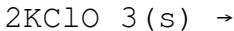
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represents a

laboratory

preparation for

oxygen gas:



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How many moles  
of  $O_2$  form if  
3.0 mol of  $KClO_3$   
are totally

...

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Answer \_\_\_\_\_ 3]

60.5 mL of HNO<sub>3</sub>

are required to

react with 25.0

mL of a 1.00 M

Barium hydroxide

solution: 2 .

HNO<sub>3</sub>(aq) +

Ba(OH)<sub>2</sub>(aq) -->

2 . H<sub>2</sub>O(s) +



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$\text{Ba}(\text{NO}_3)_2(\text{aq})$

(BALANCED) Find  
the . Molarity.  
of the nitric

acid solution .

25.0ml  $\text{Ba}(\text{OH})_2$  X

. 1L 10. 3. ml

X. 1L 1.00mol

$\text{Ba}(\text{OH})_2$ . X 2mol

$\text{HNO}_3$  1mol

$\text{Ba}(\text{OH})_2 =$

0.0500mol  $\text{HNO}_3$  .

M . . .

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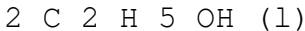
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what you are given, and what you are trying to find. 1.

Fermentation is a complex chemical process of making wine by converting glucose into ethanol and carbon dioxide:



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+ 2 CO<sub>2</sub> (g) A.  
Calculate the  
mass of ethanol  
produced if...

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correct answer

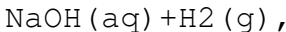
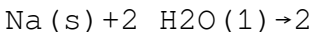
is A. H<sub>2</sub>C<sub>2</sub>O<sub>2</sub> is

the limiting

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reagent: (4.5mol  
H<sub>2</sub>C<sub>2</sub>O<sub>4</sub>) (2mol  
MnSO<sub>4</sub>/mol  
H<sub>2</sub>C<sub>2</sub>O<sub>4</sub>)

According to the  
equation 2



sodium reacts  
with water.

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