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9.1 Introduction  
to Stoichiometry

~~Chapter 9~~

~~Stoichiometry~~

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*Stoichiometry*

*Basic*

*Introduction,*

*Mole to Mole,*

*Grams to Grams,*

*Mole Ratio*

*Practice*

*Problems* **Chapter**

**9: Stoichiometry**

**examples** *Step by*

*Step*

*Stoichiometry*

*Practice*

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*to Pass*

*Chemistry* CHEM

~~103 — Chapter 9~~

~~Chemical~~

~~Equation~~

~~Calculations~~

~~(aka~~

~~Stoichiometry)~~

~~Part 1~~ **Chapter 9**

**Stoichiometry**

*Introduction to*

*Limiting*

*Reactant and*

*Page 6/41*

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Excess Reactant

Chapter 9 lesson

1 Stoichiometry

Balancing

Chemical

Equations

Practice

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Selection -

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Ideal

Stoichiometric

Calculations

Chapter 9 2 Mr C

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*Made Easy:*

*Stoichiometry*

*Tutorial Part 1*

*Limiting*

*Reactant*

*Practice Problem*

**Stoichiometry**

**Tutorial: Step**

**by Step Video +**

**review problems**

**explained |**

**Crash Chemistry**

**Academy** *High*

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*Cashtrap 2.0 -*

*Trade w/ Higher*

*Success Mole*

*Concept Tips and*

*Tricks*

~~Stoichiometry:~~

~~Converting Grams~~

~~to Grams~~

**Stoichiometry**

**Made Easy: The**

**Magic Number**

**Method** ~~Limiting~~

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~~Reactant~~  
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~~Practice Problem~~  
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~~(Advanced)~~

Avogadro's  
Number, The  
Mole, Grams,  
Atoms, Molar  
Mass

Calculations -

Introduction

Stoichiometry:

What is

Stoichiometry?

*Stoichiometry* -

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Limiting \u0026amp;

Excess Reactant,

Theoretical

\u0026amp; Percent

Yield -

Chemistry 9.2

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Calculations

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**Chapter 9 -**

**Chemical**

**Equation**

**Calculations**

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Stoichiometry)  
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(10-11) Chapter  
9 sections 3 and  
4. *Converting  
Between Grams  
and Moles MASS  
SPECTROMETER FSC*

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Waqar Ahmad

Introduction to

Moles ~~Chapter 9~~

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KEY Page 1 1.

The following

equation

represents a

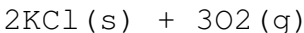
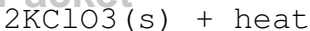
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preparation for  
oxygen gas:



How many moles

of  $\text{O}_2$  form as

3.0 mol of  $\text{KClO}_3$

are totally

consumed? 3.0

mol  $\text{KClO}_3 \times (3$

moles  $\text{O}_2) / (2$

moles  $\text{KClO}_3) =$

4.5 moles  $\text{O}_2$  2.

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~~Ch 9 Packet KEY~~

~~+ Stoichiometry~~

~~+ Mole (Unit)~~

Chapter 9

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9.3 Objectives

Describe a

method for

determining

which of two

reactants is a

limiting

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reactant.  
Calculate the  
amount in moles  
or mass in grams  
of a product,  
given the  
amounts in moles  
or masses in  
grams of two  
reactants,

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Yield

Calculations

Tutorial Link to

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Excess Reactant

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complete the

Excess Reactant

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mass of  $\text{CuO}$  to  
98.4 grams

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9.3 Objectives

Describe a  
method for  
determining  
which of two



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reactants is a  
limiting  
reactant.

Calculate the  
amount in moles  
or mass in grams  
of a product,  
given the  
amounts in moles  
or masses in  
grams of two  
reactants, one  
of which is in  
excess.;

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Distinguish

between

theoretical

yield, actual

yield, and

percentage

yield.;

Calculate

percentage

yield, given the

actual yield and

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[Stoichiometry]

Chapter Ten

[States of

Matter] Chapter

Eleven [Gases]

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review 5 thru

16. Section 3:  
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PROBLEMS Write  
the answer on  
the line to the  
left. Show all  
your work in the  
space provided.

1. 88% The

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Packet  
actual yield of  
a reaction is 22  
g and the

theoretical

yield is 25 g.

Calculate the  
percentage

yield. 2. 6.0

mol of  $N_2$  are

mixed with 12.0

mol of  $H_2$

according to the  
following

equation:  $N_2(g)$

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~~mc06se cFMsr i~~  
~~vi~~

Chapter 9  
Chemical  
Quantities 1.  
Although we  
define mass as  
the “amount of  
matter in a  
substance,” the  
units in which we  
measure mass are

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a human invention. Atoms and molecules react on an individual particle- by- particle basis, and we have to count individual particles when doing chemical calculations.

~~Chapter 9~~

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Changing

Attitudes

Shunning the

ancient Greek

approach of

logical argument

based on

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untested  
premises,  
investigators of  
the seventeenth  
century began to  
understand the  
laws of nature  
by observing,  
measuring, and  
performing  
experiments on  
the world around  
them.

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~~Collaborative~~

Unit 6 Packet -

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Vocab Assignment

Due: Tuesday,

Dec. 2 nd

Problem Set Due:

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Thursday, Dec. 9  
th Test Date:

Friday, Dec. 10

th VOCABULARY

Assignment:

stoichiometry

percentage yield

mole ratio mass-

mass problem

limiting reagent

excess reagent

OBJECTIVES: •

~~Unit 6~~

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Chapter 12

Stoichiometry

Packet Answers

In Example

12.2.1 and

Example 12.2.2,

the identity of

the limiting

reactant has

been apparent:

[Au(CN)<sub>2</sub>]<sup>-</sup>,  
LaCl<sub>3</sub>, ethanol,

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and para-  
nitrophenol.  
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