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equations. (Coefficients equal to one (1) do not need to be shown in your answers). (a) $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$. (b) $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. (c) $2\text{FeBr}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 1\text{Fe}_2(\text{SO}_4)_3 + 6\text{HBr}$ (d) $1\text{C}_4\text{H}_6\text{O}_3 + 1\text{H}_2\text{O} \rightarrow 2\text{C}_2\text{H}_4\text{O}_2$.

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~~Answers Key~~
equations and indicate the type of reaction taking place: 1) $3. \text{NaBr} + 1. \text{H}_3\text{PO}_4 \rightarrow 1. \text{Na}_3\text{PO}_4 + 3. \text{HBr}$ Type of reaction: double displacement. 2) $3. \text{Ca}(\text{OH})_2 + 1. \text{Al}_2(\text{SO}_4)_3 \rightarrow 3. \text{CaSO}_4 + 2. \text{Al}(\text{OH})_3$ Type of reaction: double displacement. 3) $3. \text{Mg} + 1. \text{Fe}_2\text{O}_3 \rightarrow 2. \text{Fe} + 3. \text{MgO}$ Type of reaction: single displacement. 4)

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